

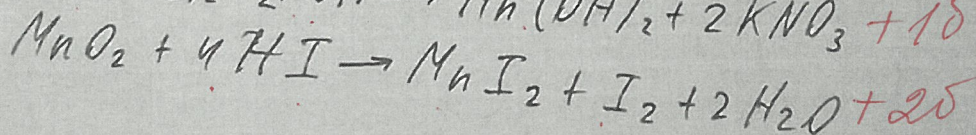
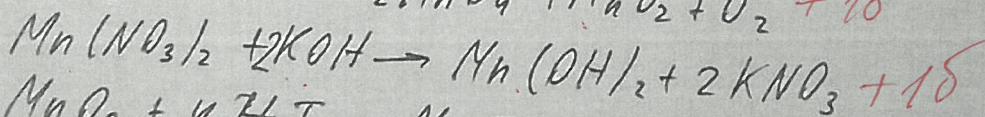
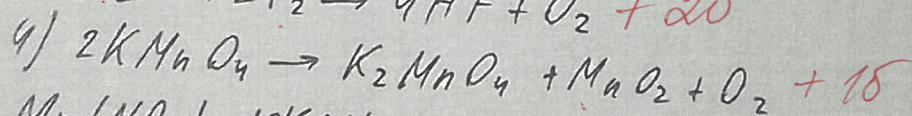
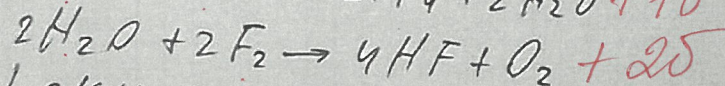
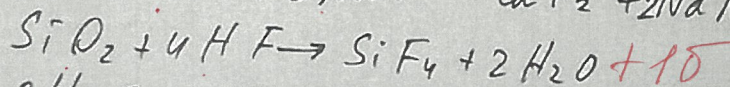
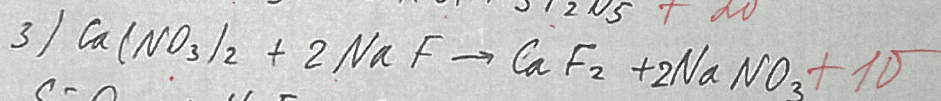
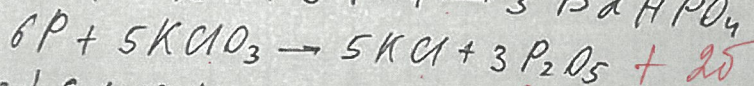
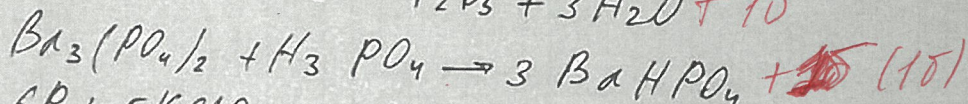
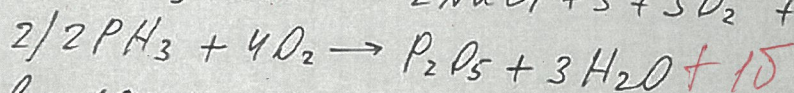
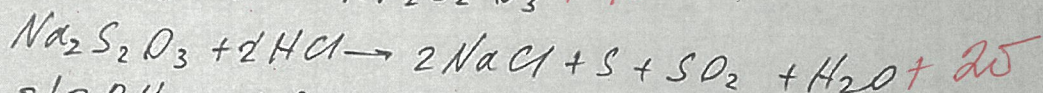
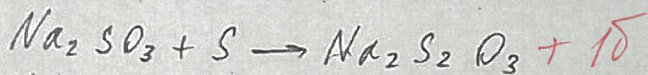
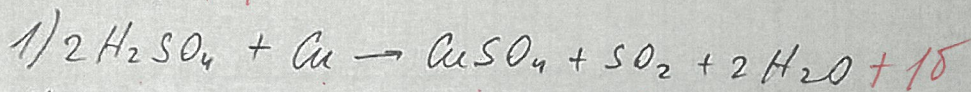
NS.

$$A-S + 15$$

$$D-P + 15$$

$$E-F + 15$$

$$G-Mn + 15$$



$$\Sigma 205 \text{ (задание NS)}$$

1	2	3	4	5	6	Σ
2	14,5	125	—	20	205	68,5

№1.

Э-N^+
 N_2O и NO_2 + } 25

~~CO и CO_2~~

~~CO - бесцветный газ без запаха, ядовит~~

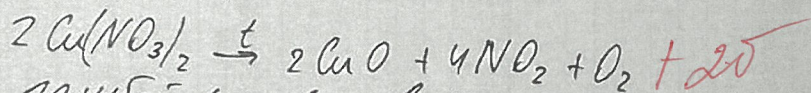
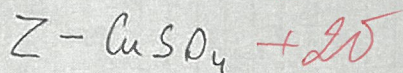
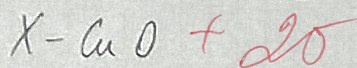
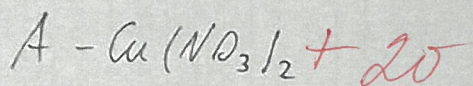
~~CO_2 - бесцветный газ без запаха, используется для подкормки мизин растений~~

~~Fe_2O_3 и Fe_3O_4~~

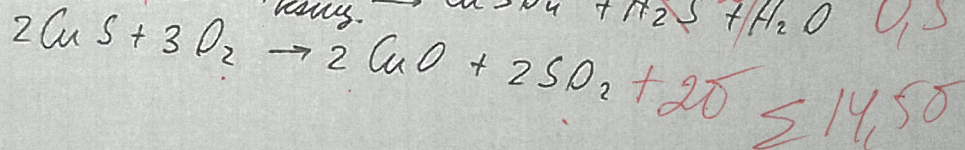
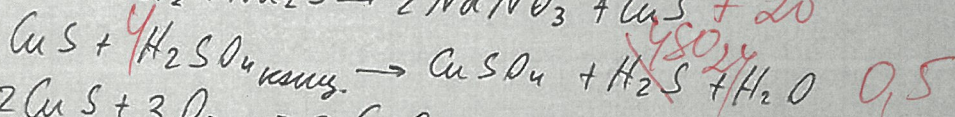
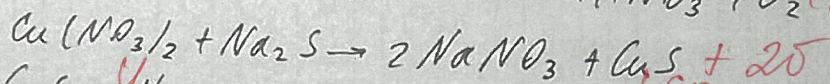
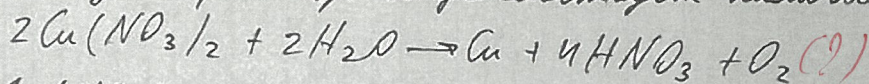
~~Fe_2O_3 - красно-коричневое твёрдое вещество, растворимо в кислотах~~

~~Fe_3O_4 - чёрное твёрдое вещество, растворимо в кислотах~~

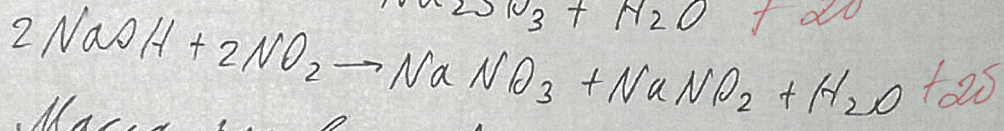
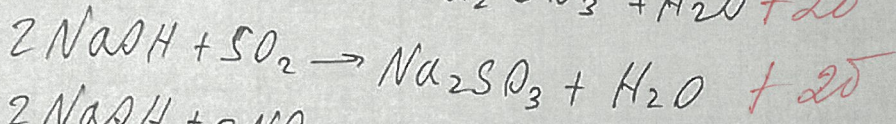
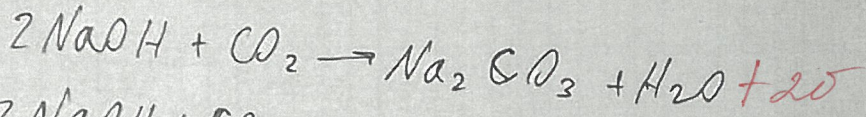
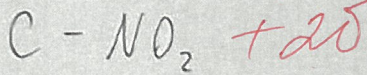
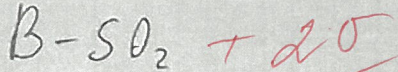
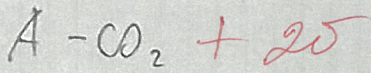
N2



любой раствор свидетельствует наличие меди



N3



Масса раствора NaOH увеличится только на 25 г, в остальных реакциях не изменится, поэтому прирост будет 0.

$$\Sigma 125$$

№ 6.

$$M(C_{10}H_8) = 12 \cdot 10 + 8 = 128 \text{ г/моль} +$$

$$n(C_{10}H_8) = \frac{m}{M} = \frac{64}{128} = 0,5 \text{ моль} +$$

$$\mu = \frac{n \text{ растворен. в-ва}}{m \text{ растворителя}}$$

$$\mu = \frac{0,5}{0,25} = 2 \text{ моль/кг} (+) \quad 60$$

$$T_{\text{кристал.}} = T_{\text{пл.}} - \Delta T_{\text{зам.}}$$

$$\Delta T_{\text{зам.}} \text{ раствора } C_{10}H_8 \text{ в } C_6H_6 = 2,57 \cdot 2 = 5,14^\circ\text{C} (+)$$

$$T_{\text{кристал.}} (C_{10}H_8 \text{ в } C_6H_6) = 5,5 - 5,14 = 0,36^\circ\text{C}$$

$$0,36^\circ\text{C} + 273,15 = 273,51 \text{ K} (+)$$

$$\Delta T_{\text{зам.}} \text{ раствора } C_{10}H_8 \text{ в } CHCl_3 = 3,88 \cdot 2 = 7,76^\circ\text{C} (+)$$

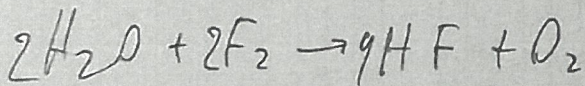
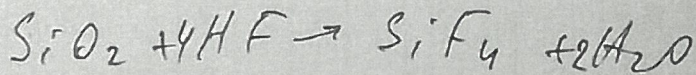
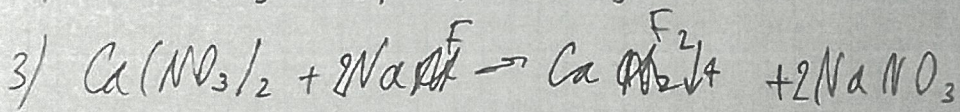
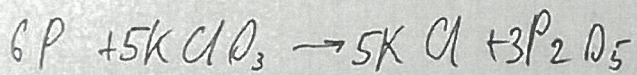
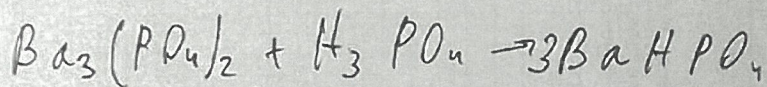
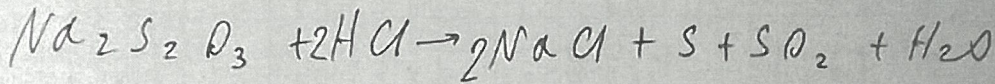
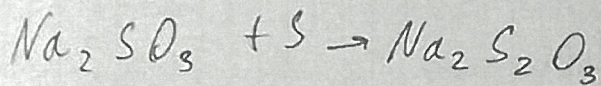
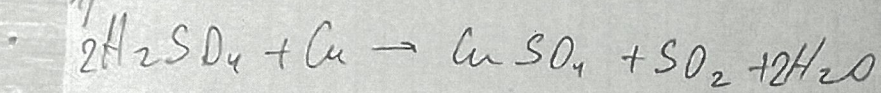
$$T_{\text{кристал.}} (C_{10}H_8 \text{ в } CHCl_3) = -63,5 - 7,76 = -71,26^\circ\text{C}$$

$$-71,26^\circ\text{C} + 273,15 = 201,89 \text{ K} (+)$$

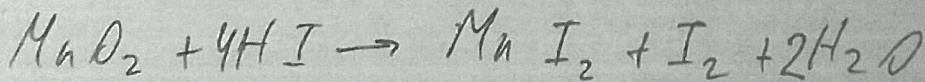
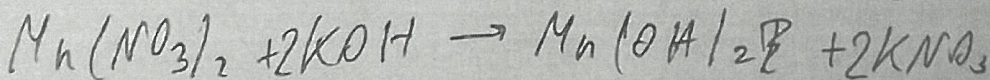
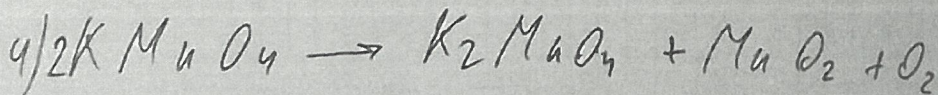
Σ 205.

(5)

$\therefore$ A-S D-P ^{NS} E-F G-Mn
 1/



нобоор



~1.

N_2O и NO_2 Э-N

Fe_2O_3 и Fe_3O_4

нобоор

красно-коричневый, бл. б-го
разл. с углеродом

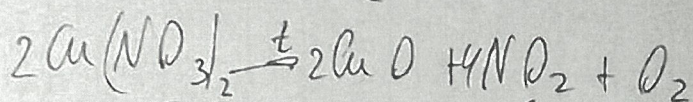
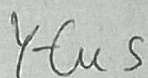
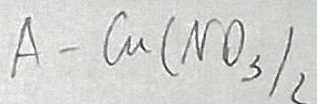
темный, бл. б-го

CO CO_2 - бесцвет. газ без запаха

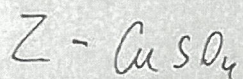
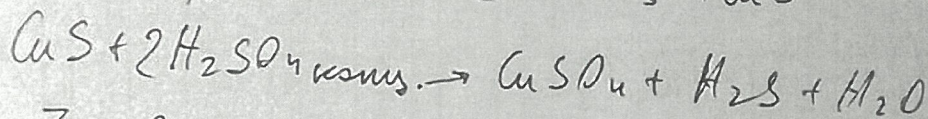
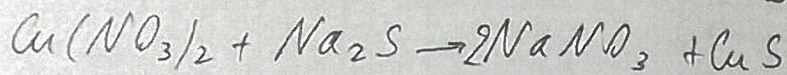
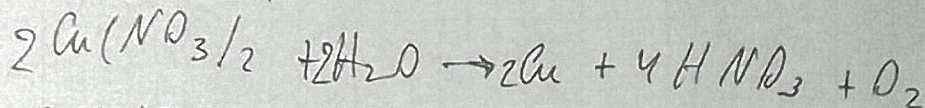
бесцвет.
газ без
запаха,
является

исп. для дыхания
можно раскислить

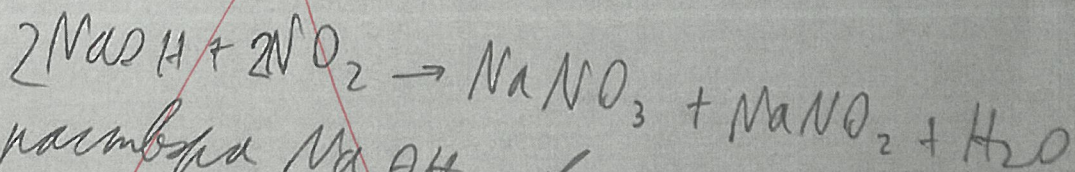
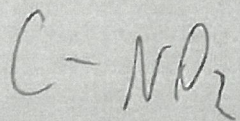
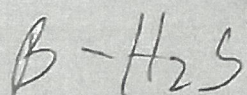
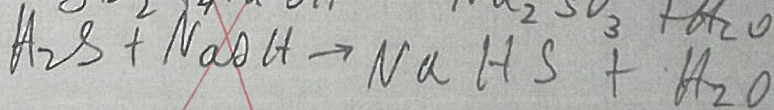
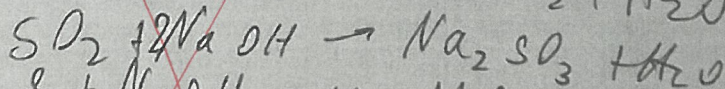
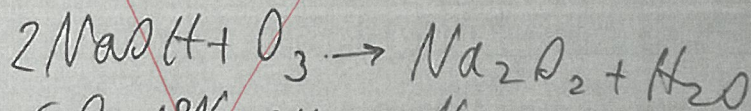
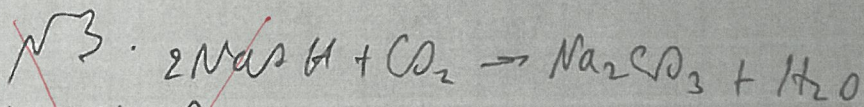
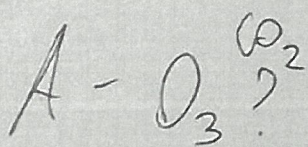
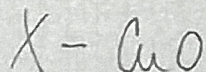
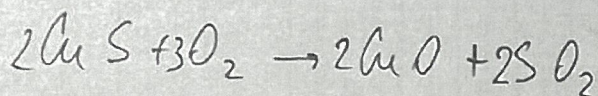
№ 2



изучили реакцию свинцовым и медью



повтор



масса раствора NaOH увеличивается
на 10% в результате, в результате реакции не идет,
потому что будет 0

№6.

X325803

$$M(C_{10}H_8) = 12 \cdot 10 + 8 = 128 \text{ г/моль}$$

$$n(C_{10}H_8) = \frac{m}{M} = \frac{64}{128} = 0,5 \text{ моль}$$

$$\mu = \frac{\text{криоскоп. д-во}}{m \text{ растворителя}}$$

$$\mu = \frac{9,5}{4,75} = 2 \text{ моль/кг}$$

$$T_{\text{крист. изм.}} = T_{\text{наблев.}} - \Delta T_{\text{зам.}}$$

$$\Delta T_{\text{зам.}} \text{ (C}_{10}\text{H}_8 \text{ в C}_6\text{H}_6) = 2,57 \cdot 2 = 5,14^\circ\text{C}$$

$$T_{\text{крист.}} = 5,5 - 5,14 = 0,36^\circ\text{C}$$

$$0,36^\circ\text{C} + 273,15 = 273,51 \text{ K}$$

или

$$\Delta T_{\text{зам.}} \text{ (C}_{10}\text{H}_8 \text{ в CHCl}_3) = 3,88 \cdot 2 = 7,76^\circ\text{C}$$

$$T_{\text{крист.}} (C_{10}H_8 \text{ в CHCl}_3) = -63,5 - 7,76 = -71,26^\circ\text{C}$$

$$-71,26 + 273,15 = 201,89 \text{ K}$$

повтор